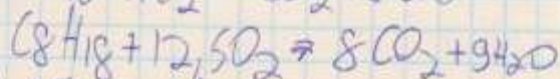
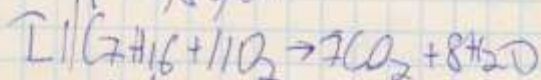


$$2) \text{C}_8\text{H}_{18}$$

$$\rho = 100 \text{ kg/m}^3$$

$$V = 10 \text{ m}^3$$

$$\rho = 0,82 \text{ g/cm}^3$$



$$2) m(\text{benzina}) = 10000 \cdot 0,82 \text{ g/cm}^3 = 8000 \text{ g}$$

$$3) m(\text{C}_7\text{H}_{16}) = 8000 \cdot 0,2 = 1600 \text{ g}; n(\text{C}_7\text{H}_{16}) = \frac{1600}{100} = 16 \text{ mola}$$

$$m(\text{C}_8\text{H}_{18}) = 8000 \cdot 0,8 = 6400 \text{ g}; n(\text{C}_8\text{H}_{18}) = \frac{6400}{114} =$$

$$4) n(\text{O}_2) = 16 \cdot 11 = 176 \text{ mola}$$

$$n(\text{O}_2) = 56,14 \cdot 12,5 = 701,75 \text{ mola}$$

$$n(\text{CO}_2) = \dots \text{ mola} \cdot V(\text{O}_2)_{\text{I,II}} = 1966,6 \text{ l}$$

$$5) V(\text{CO}_2) = \frac{V(\text{O}_2)}{p(\text{O}_2)} = \frac{1966,6}{0,2} = 9833,3 \text{ l} (9,833 \text{ m}^3)$$

$$6) V(\text{CO}_2)$$

$$n(\text{CO}_2) = 16 \cdot 7 = 112 \text{ mola}$$

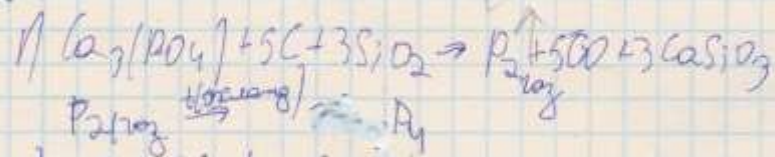
$$n(\text{CO}_2) = 56,14 \cdot 8 = 449,12 \text{ mola}$$

$$n(\text{CO}_2)_{\text{I,II}} = 56,12 \text{ mola}; V(\text{CO}_2)_{\text{I,II}} = 56,12 \cdot 22,4 = 1256,9 \text{ l} (1,2569 \text{ m}^3)$$

120

Oxidation: 12,57 wt%

10



60

